## **Thinking Electronegatively HW**

Read and outline (Cornell style) **Section 9.5** in your chemistry textbook. Then answer the following assessment questions:

- 1. Define electronegativity. What is the periodic table trend for electronegativity?
- 2. How is the electronegativity difference used in determining the type of bond that occurs between two atoms?
- 3. Describe a polar covalent bond. Describe a nonpolar covalent bond. Describe an ionic bond.
- 4. What is a polar molecule?
- 5. Predict the type of bond that will form between the following atoms:
  - a. H and S
  - b. Cand H
  - c. Na and S
- 6. Complete the following chart:

Molecule	Lewis Dot Structure (use Octet Rule)	Structural Formula (use HONC 1234)	Symmetrical shape?	All surround- ing atoms identical?	Polar or Non- polar?	Can we smell it?	Dissolves in water?	Type of IMF's? (Intermolecular Forces)
SiCl <sub>4</sub>								
H <sub>2</sub> S								
SiH <sub>2</sub> O								
SiO <sub>2</sub>								
NH <sub>3</sub>								